**Karan Arora**  **R.L. Institute M: 9416974837**

**Max Time : 1 hr** **Class = 11th Chemistry Test Max Marks : 30**

**Redox**

(Oxidation Number , Acidic and Basic Balancing)

1. Balance the following equations by acidic medium [ 4 x 2 = 8 ]
2. Sn (s) + (aq) + H + (aq) Sn2+ (aq) + (aq) + H2O (l).
3. (aq) + C2H4O (g) → Cr3+ (aq) + C2H4O2 (aq).
4. Cu (aq) + (aq) Cu2+ (aq) + NO2 (g).
5. (aq) + Fe2+ (aq) → Mn2+ (aq) + Fe3+ (aq)
6. Balance the following equations by Basic medium [ 2 x 2 = 4 ]
7. Cr (s) + (aq) Cr(OH)3 (s) + (aq)
8. Zn (s) + (aq) Zn2+ (aq) + (aq)
9. Find the oxidation number of the following : [ 1 x 8 = 8]

(a) C2H6 (b) (NH4)2SO4

(c) KClO4 (d) PbSO4

(e) SiH4 (f) KMnO4

(g) Cr (h) H4P2

1. Identify the oxidizing , reducing agent and which will act as both oxidizing and reducing, of the following. [ 1 x 10 = 10]

(a) HNO3 (b) H2SO4

(c) HNO2 (d) H2S

(e) NF3 (f) H3PO4

(g) H3PO2 (h) H3PO3

(i) (j) SO2